

## PAPER DETAILS

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## THE EFFECT OF GRAPHENE OXIDE ON THE STRUCTURAL AND ELECTRICAL PROPERTIES OF YTTRIUM FERRITE BASED NANOPOWDERS

*Handan AYDIN*

*In this study, Yttrium Ferrite based nano electroceramics were prepared by the doping at various ratios of graphene oxide by the hydrothermal method. The obtained doped and undoped samples were characterized by XRD, SEM, FTIR, DTA, UV-VIS-NIR spectroscopy, temperature dependent electrical resistance changes and dielectric measurements. From the XRD results, it is observed that the obtained samples are polycrystalline with an Orthorhombic structure. The crystal parameters are changed with the doping of graphene oxide. FTIR spectra indicate that the graphene oxide changes the peak intensities. This confirms the incorporation of the graphene oxide into the ceramic structure. The doping of the graphene oxide changes the optical band gaps of the samples. The dielectric constants of the samples are changed with the graphene oxide. The DC electrical conductivity measurements confirm the semiconducting behavior of the samples. The obtained results indicate that the electronic and nanostructure of the synthesized electroceramics can be controlled by the doping of graphene oxide for electronic and energy applications.*


**Keywords:** *Hydrothermal method, Graphene oxide, Hummers method, Nano electroceramic, structural properties.*

### 1. Introduction

Over the last years, research on the nano-structured electroceramics has been increasing. These materials have played an important role in advanced technology applications [1] such microwave industries [2], electronics [2], data storage devices [2, 3], gas-sensitive sensors [4], digital disk recording[2] and actuator applications [2]. Efforts have been made for the improving of specific properties and potential applications [5] of nano electroceramics  $ABO_3$  type (where A is a rare or alkaline earth metal and B is a first-row transition metal) [6, 7]. The materials in this formula are often referred to as rare earth orthoferrites [8]. Orthoferrites have attracted attention due to their a wide range of Ferro-, piezo-, pyroelectrical properties and electro-optical effects[7]. Yttrium ferrite ( $YFeO_3$ ) is a member of this class. It has perovskite structure with an orthorhombic unit cell [5, 8, 9] and a p-type semiconductor ( $E_g=2.58$  eV) [4].

Recently, the studies on the structural and electrical properties of yttrium ferrite are increasing. In this context, the scientists have investigated and concentrated zoom in on improving the properties of Yttrium Ferrite with a different dopant and various element.

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In recent years, to produce these materials have been used different synthesis process such as solvothermal [2], sol-gel [2, 5, 6, 8-11], , hydrothermal method [6, 8, 9], coprecipitation [8, 9], solid state reaction method [5, 6, 9, 10], self-propagating combustion synthesis [5], alkoxide method [5, 10], , pulsed laser deposition [5, 10], microwave-assisted [5, 9, 10], sonochemical synthesis [2, 5, 10]. Among these production methods, hydrothermal synthesis can be preferred due to its advantages such low temperature, control of particle size, without any calcination or milling step and not purity phase [12]. In this scientific study, pure and Graphene oxide doped Yttrium Ferrite nano electro-ceramics were synthesized by the Hydrothermal Process. The structural, morphological, chemical and physical electrical properties of the obtained samples were investigated.

## 2. Experimental Details

### 2.1. Synthesis of Graphene Oxide

Graphene oxide was prepared by a modified Hummers method. In this method, 1 g graphite powder, 0.5 g sodium nitrate and 23 ml sulfuric acid were mixed into a beaker and exposed stirring for 30 min in an ice bath. Then 3 g potassium permanganate was slowly added to mixture solution by magnetic stirring through 6 hours. Subsequently, after the solution was kept in the ultrasonic machine for 2 hours and 100 ml distilled water and 2ml H<sub>2</sub>O<sub>2</sub> (30%) were added by using a magnetic stirrer. Finally, the solution was washed with de-ionized water and filtered until pH value reached to 7.

### 2.2. Synthesis of YFeO<sub>3</sub> via Hydrothermal Method

Synthesis of YFeO<sub>3</sub> nano electroceramics were carried out by hydrothermal route. The precursor solution was prepared using Yttrium (III) nitrate hexahydrate (N<sub>3</sub>O<sub>9</sub>Y) (Abcr, %99,9), Iron (III) Chloride Hexahydrate (FeCl<sub>3</sub>.6H<sub>2</sub>O) (Merck, %99) and potassium hydroxide (KOH) (Merck, %85) as mineralizer and de-ionized water as a reaction medium. For the production process of the samples, the schematic diagram about with Hydrothermal method is given in Fig. 1. In this reaction procedure, four solutions containing N<sub>3</sub>O<sub>9</sub>Y, 5.5 mmol FeCl<sub>3</sub>.6H<sub>2</sub>O, and 9 mol KOH were prepared. 6.1 mmol of N<sub>3</sub>O<sub>9</sub>Y was added into 55mL of de-ionized water with constant magnetic stirring followed by the addition of 6.1 mmol of FeCl<sub>3</sub>.6H<sub>2</sub>O. The solutions were stirred at room temperature for 1 h to obtain a homogeneous solution. Thereafter, KOH was slowly added to the above Y-Fe solution to co-precipitate. The solution was kept at room temperature °C for 2 h under stirring. After, each graphene oxide solution was added to the precursor solution. The final mixture is here after called LFO, LFO20, LFO50 and LFO100, respectively. The suspension solution was poured into full-automatically hydrothermal system (Fytronix HT-100) for the hydrothermal process. The process was carried out for 6 h at temperature of 200 °C. After completion of the hydrothermal reaction, the autoclave was cooled naturally to the room temperature. The products were filtered and dried at room temperature for 24 h. The obtained powders were formed by hydraulic press machine under a uniaxial pressure of 10 MPa at room temperature.

### 2.3. Characterization of Samples

The crystalline phase of the products was recognized using a BRUKER ADVANCE D8 X-ray diffractometer with Cu K $\alpha$  radiation (  $\lambda = 1.5406 \text{ \AA}$ ) in the 2 $\theta$  range from 20 to 80° with 0.02 °/min. at room temperature. The morphology of nano electro-ceramics were investigated with the a High-Resolution Scanning electron microscope (JEOL JSM-7001F). High-Resolution Transmission electron microscope (JEOL JEM-2100F) was performed at an accelerating voltage of 200 kV, by placing the powder on a copper grid to observe the microstructure of the powders. The optical spectrum was recorded by a SHIMADZU 3600 UV-VIS-NIR spectrophotometer in the wavelength range of 200–1400

nm at room temperature. The thermal analysis of the powders were determined by SHIMADZU DTG-60AH from room temperature to 900 °C in Argon ambient at a scan rate of 10 °C/min. FTIR spectra studies were made with Thermo Scientific Nicolet IS5 mark Fourier transform infrared spectrometer with an ID5 ATR sample holder in the wavenumber range from 400 to 4000 cm<sup>-1</sup>. The dielectric and alternating current conductivity properties were investigated using a HIOKI 3532-50 LCR HITESTER at room temperature. The ferroelectric properties were studied using Radiant Precision Premier II Technology. The electrical conductivity of the samples was measured as a function of temperature using a KEITHLEY 6517A electrometer.

### 3. Results and Discussion

#### 3.1. Structural and morphological characterization of Produced Nano Electroceramics

The XRD patterns of pure and GO doped YFeO<sub>3</sub> nanostructure powder samples are shown in Fig.1. The observed all peaks in Fig.1. revealed that YFeO<sub>3</sub> are polycrystalline structure with the orthorhombic structure. The crystalline phases were identified by reference to the Joint Committee on Powder Diffraction Standards (JCPDS 01-073-1345) files. XRD pattern of the samples were appeared at 2θ values, 33.462°, 35.956° and 54.361° corresponding to (200), (120) and (310) reflections, respectively. As seen in the XRD peaks, the intensity of peaks is changed with increasing Graphene oxide doping. This indicates that Graphene oxide is incorporated into sites of YFeO<sub>3</sub>.

The lattice parameters (a, b and c) of the orthorhombic perovskite structure of YFeO<sub>3</sub> at various graphene oxide contain were determined according to the following relations [6, 13,30]:

$$\frac{1}{d_{hkl}^2} = \frac{h^2}{a^2} + \frac{k^2}{b^2} + \frac{l^2}{c^2} \quad (1)$$

Where (h, k, l) are the indices of every peak of crystallographic plane obtained from XRD patterns. d is the interplanar distance, (a, b, c) are the lattice constants. Average crystallite sizes were estimated using the Debye-Scherrer equation [1, 14-19,31]:

$$D = \frac{0.9 \lambda}{\beta \cdot \cos \theta} \quad (2)$$

The crystallite parameters of the unit cells of the samples are given in Table 1. As seen in Table 1, with the addition of graphene oxide, the crystallite size was changed for (200) crystal plane.

**Table 1.** The calculated values of the crystallite parameters of the unit cells of the samples.

	2 θ (°)	Crystal Structure	FWHM	Crystallite Size (nm)	d (nm)
<b>YFO</b>	33,462	Orthorhombic	0,304	27,30	0,2675
<b>YFO20</b>	33,484	Orthorhombic	0,330	25,15	0,2674
<b>YFO50</b>	33,312	Orthorhombic	0,225	36,87	0,2687
<b>YFO100</b>	33,352	Orthorhombic	0,209	39,69	0,2684

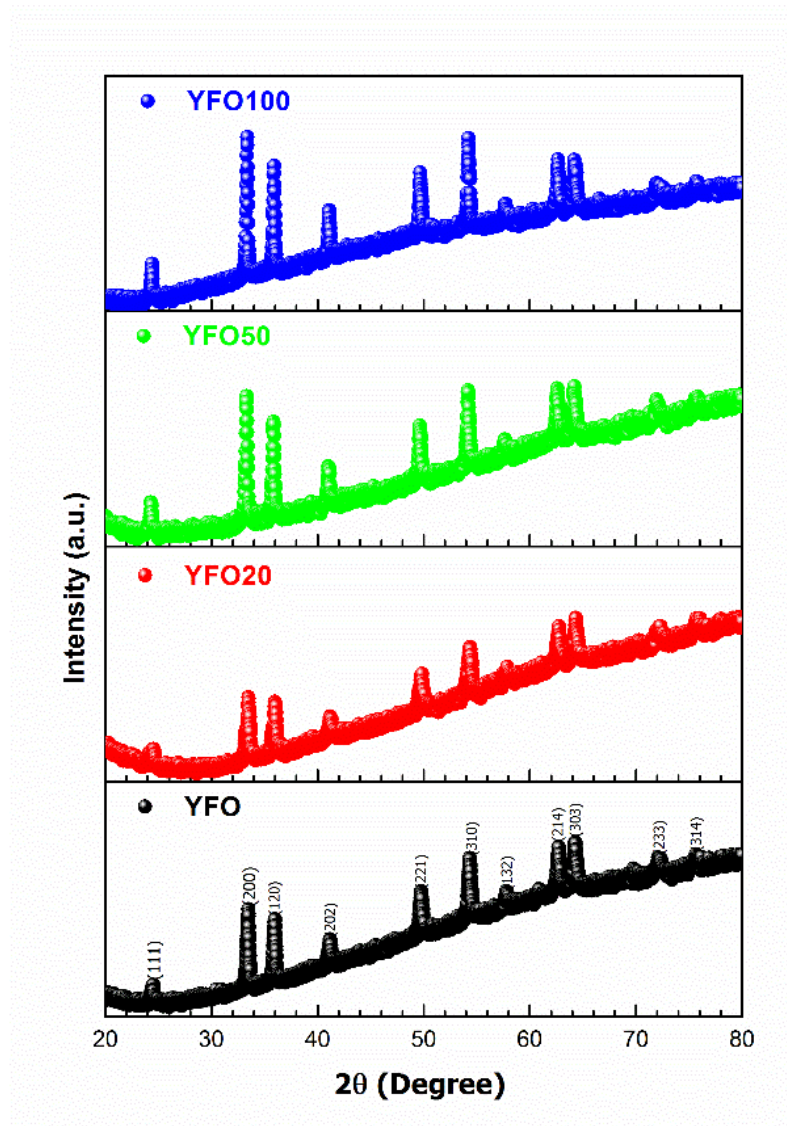
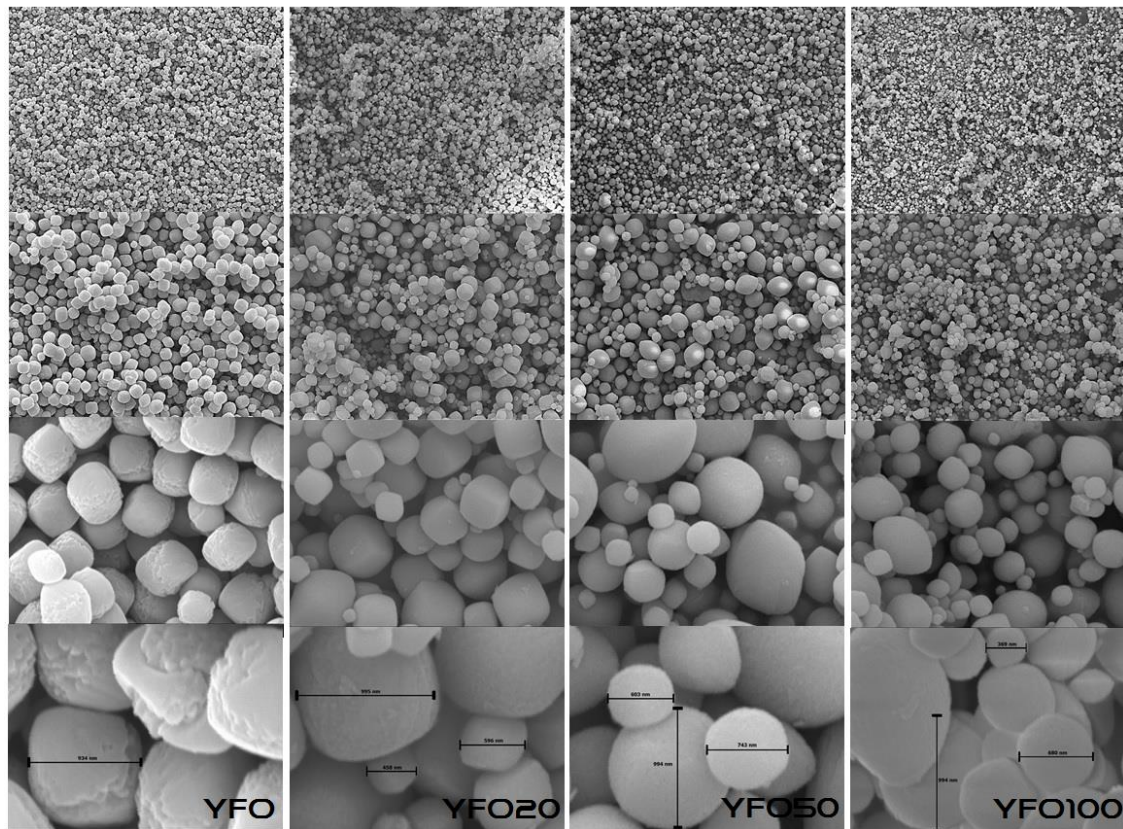
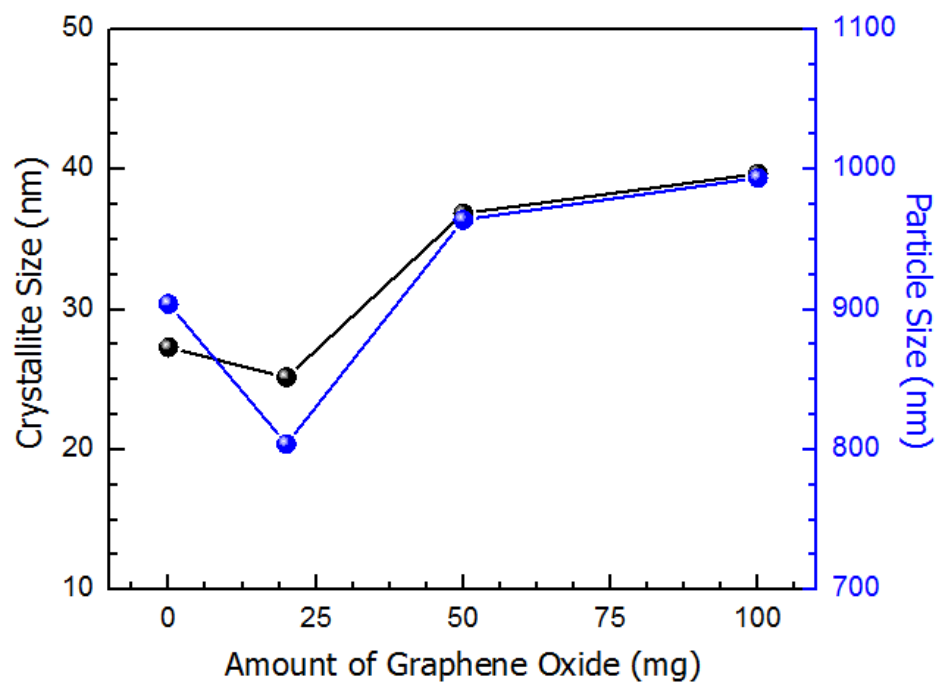


Figure 1. XRD spectrum of Samples

In order to determine the surface morphologies of undoped and graphene oxide-doped  $\text{YFeO}_3$  samples prepared by using the hydrothermal method they were analyzed with Scanning Electron Microscope (SEM). Fig.2. shows SEM images of the prepared samples. When the SEM images in Fig.2. were examined, it was seen that the grains on the surface of the sample were large and small, were randomly oriented, and were not distributed homogeneously and there were partial accumulations formed by grain stacks. It could be clearly seen that the samples formed in composite structure made of nano-sized grains and there were morphological differences in surface structure. The surface structure of the samples was generally composed of grains with similar size and morphology. It was very obvious that the surface morphologies of the samples changed with the graphene oxide dopant [32]. The change of crystallite size and particle size with the amount of graphene oxide are portrayed in Fig.3.

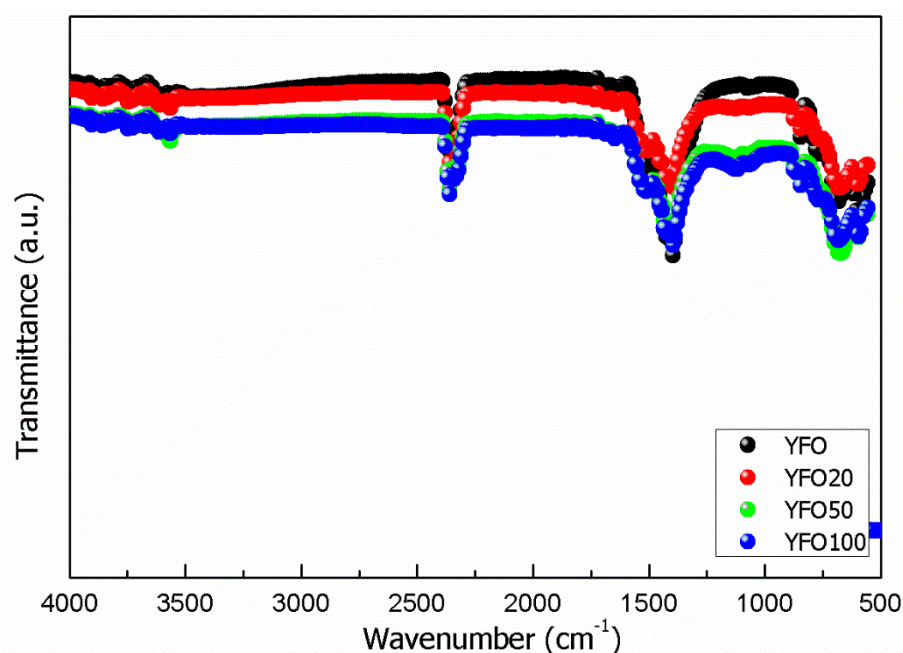


**Figure 2.** Scanning Electron Microscopy photographs of the Yttrium Ferrite nanopowders



**Figure 3.** The change of crystallite size and particle size with the amount of graphene oxide

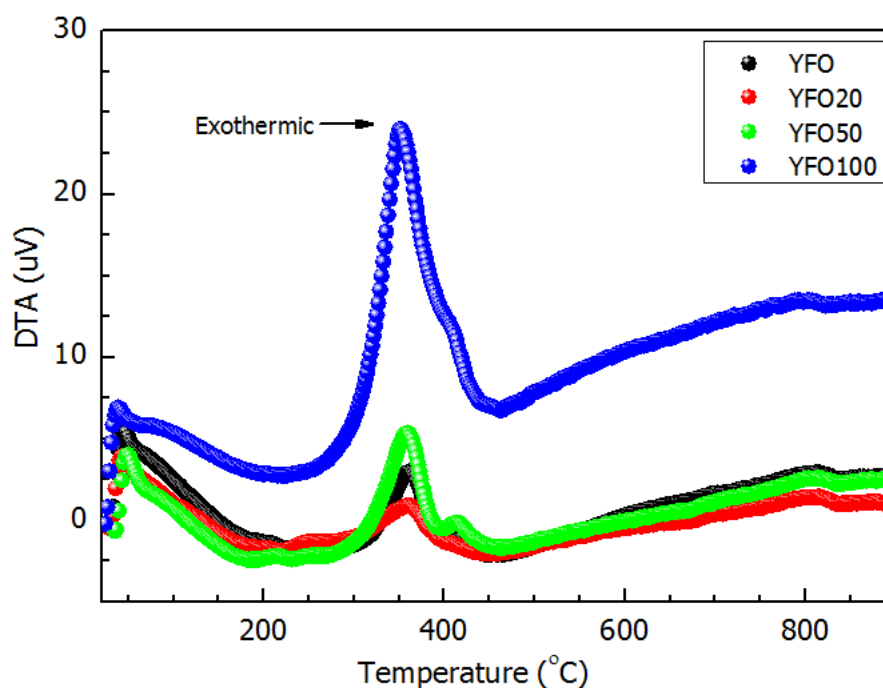




**Figure 4** FT-IR spectra of synthesized yttrium ferrite nano powders

The FTIR spectra of the samples are shown in Fig. 4., and the as-detected bands can be explained as follows. The band detected at  $1507\text{ cm}^{-1}$  is attributed to the bending mode of  $\text{H}_2\text{O}$  band. The characteristic band of  $\text{CO}_2$  is detected at  $2364\text{ cm}^{-1}$ . The band at  $1397\text{ cm}^{-1}$  is due to the  $\text{C}=\text{O}$  bond. The band observed at  $1320\text{--}1360\text{ cm}^{-1}$  is related to the  $\text{NO}_3^-$  stress vibration mode. The band at  $592\text{ cm}^{-1}$  is related to the bending vibration mode of  $\text{Fe}-\text{O}$  bond.

Fig. 5. shows both DTA curves in the range from room temperature to  $900^\circ\text{C}$  for each sample. There are one exothermic peaks in the DTA thermograms of all the samples. At the temperatures below the this peak, the weight losses are due to the elimination of the adsorbed water from the samples.

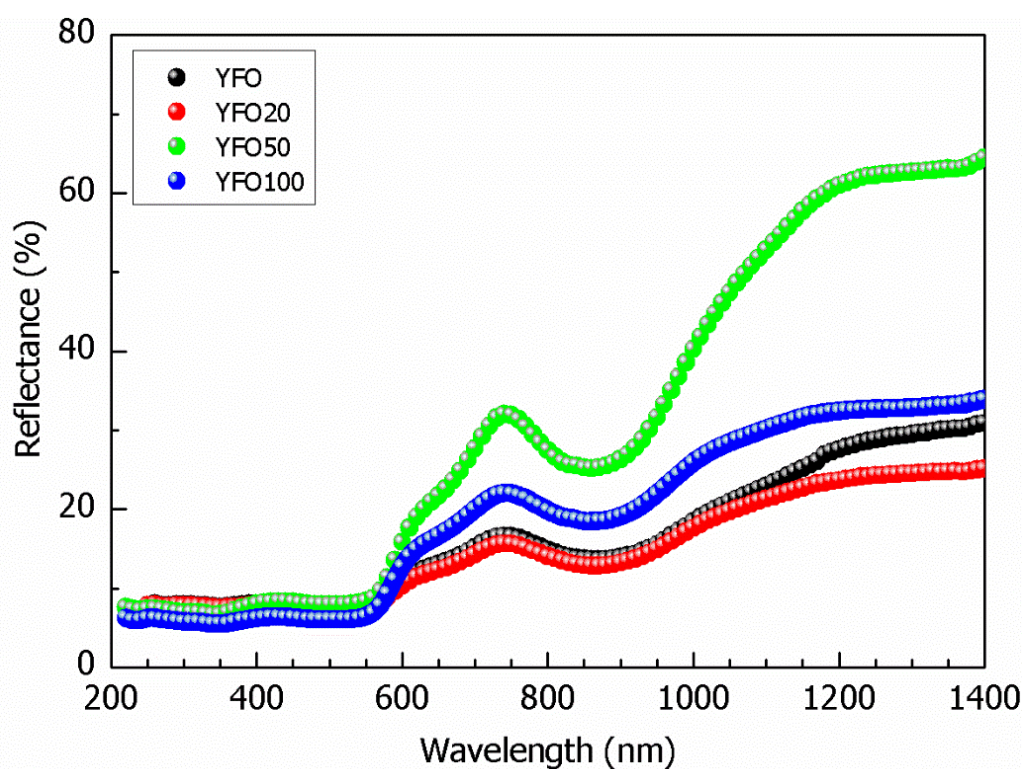


**Figure 5.** DTA curves for the nanopowders

For the peaks for each sample, these losses may be due to the elimination of some oxygen-containing functional groups. While the peak is detected at 358, 352, 355 and 348 °C for YFO, YFO20, YFO and YFO100, respectively. As can be easily seen from DTA thermograms, these peaks detected at the above-mentioned temperatures are due to the weight loss for each sample. These findings mean that the amount of the additive affects significantly the thermal stability of the GO.

### 3.2. Optical characterization

Reflection spectra of undoped and graphene oxide doped  $\text{YFeO}_3$  samples are presented in Fig. 6. As shown in Fig. 6, the reflection curves of the samples show a rapid decrease in a wavelength of about 600 nm. This peak varies depending on graphene oxide adding. This confirms that the optical band gaps of the films change with graphene oxide doping. With graphene oxide doping, the reflectance values of the samples increase compared to the undoped  $\text{YFeO}_3$  sample. The change in optical reflectance may be due to the morphological change in the samples because the aspect ratio of the crystallites varies depending on the sample surfaces. The results obtained confirm the results obtained in the SEM photographs. The sample with the highest reflectance value is the undoped YFO50 sample. Besides, the reflectance values decreased at shorter wavelengths. The reason for this may be the reflection of photons due to the decrease in their energy and thus less interaction with electrons, atoms, or crystal molecules.



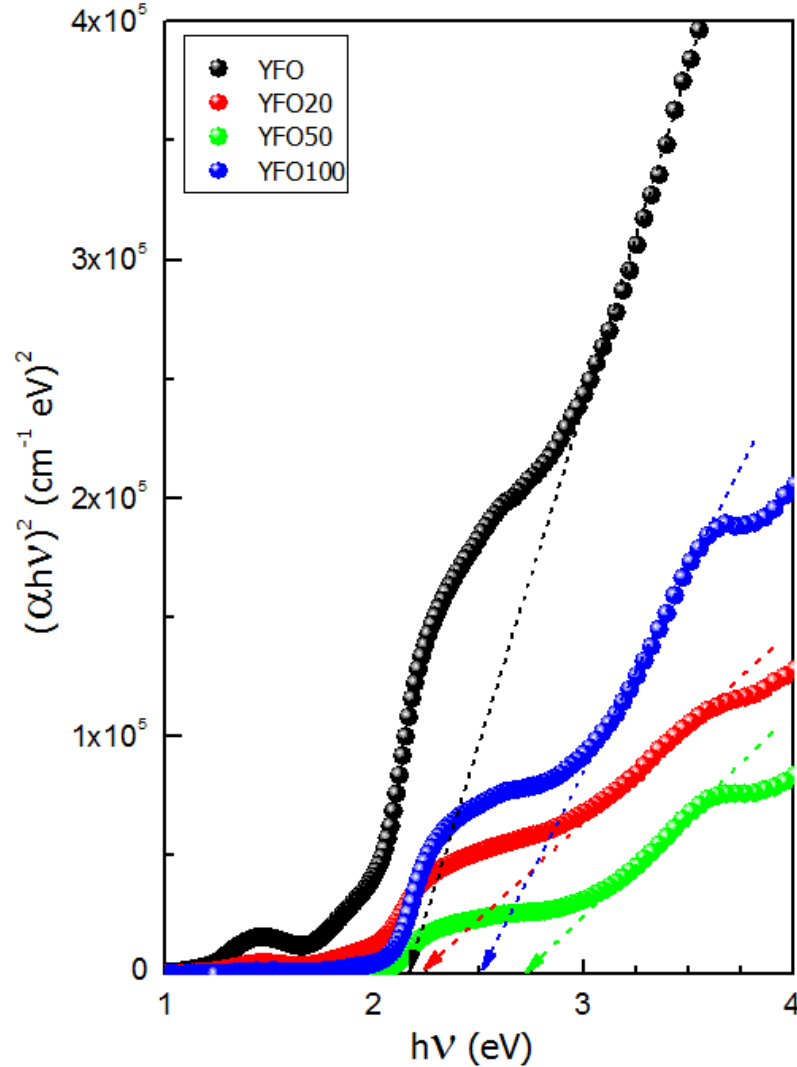
**Figure 6.** The reflectance spectra of samples as a function of wavelength

The optical band gap ( $E_g$ ) had a direct effect on the optical properties of the semi-conductive samples. The diffuse reflectance spectra given in the Fig. 6. were used to determine the band gaps ( $E_g$ ) of the undoped and graphene oxide doped  $\text{YFeO}_3$  nano powders by using the Kubelka-Munk function. Kubelka-Munk function [18] can be defined as follows [1, 13, 16, 17, 19-24,33-35] :



$$\left(\frac{F(R)hv}{t}\right) = A \cdot (hv - E_g)^n \quad (3)$$

Where  $R$  is diffuse reflectance;  $F(R)$  is Kubelka-Munk function corresponding to the absorption value;  $t$  is the thickness of the sample;  $A$  is an energy-independent constant,  $hv$  is photon energy, and  $E_g$  is band gap and  $n$  is  $\frac{1}{2}$  for direct transitions.

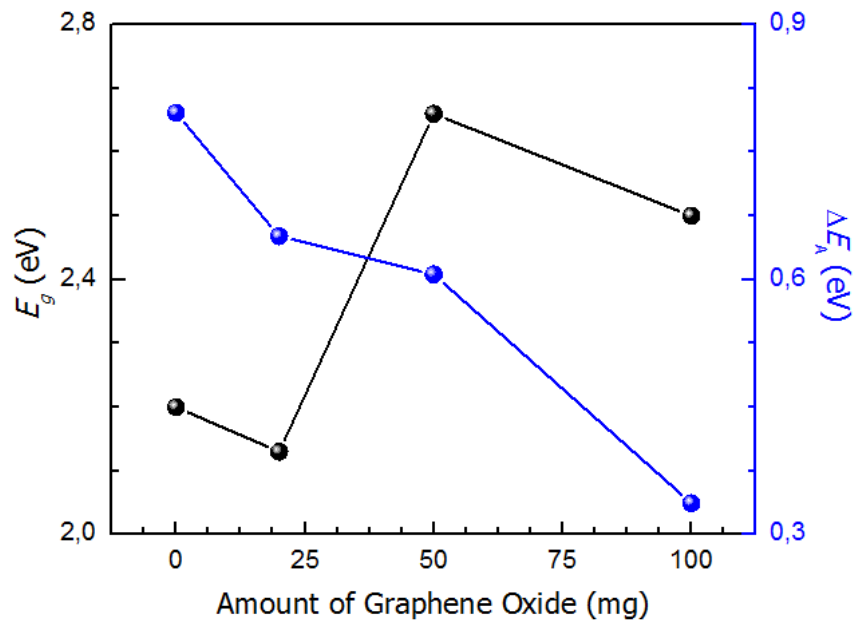


**Figure 7.** The plot of Kubelka–Munk relation for the yttrium ferrite nano electroceramics

The  $hv-(\alpha hv)^2$  graphs of the samples were drawn and the point where the line opposite to the linear part of the graph intersected the axis  $hv$  at 0 was determined. Fig.7. shows the  $hv-(\alpha hv)^2$  graphs of the samples. The values of the band gap ( $E_g$ ) calculated with the help of these curves were shown in the Table 2. As can be seen from the Table 2,  $E_g$  values of the samples increased as the atomic rate of the doped Fe increased. The reason of the increase in the band gaps of the samples was the increase in the crystal size. Since the optical properties are were affected by the changes in the size of the nano-structure. The change in the band gap with the doping process is known as "Burstein-Moss" effect. The energy band gap is in the visible region and hence the  $YFeO_3$  nanopowders in the current study are a good candidate for optical applications[25]. The variation of band gap and activation energy with the amount of graphene oxide are presented in Fig. 8.

**Table 2.** Optical band gap ( $E_g$ ) and the activation energy ( $\Delta E_A$ ) of the nanopowders

	$E_g$ (eV)	Activation Energy ( $\Delta E_A$ ) (eV)	$\sigma_{dc}$ (S/cm) (30°C)
<b>YF0</b>	2,20	0,796	$1,09 \times 10^{-6}$
<b>YFO20</b>	2,13	0,651	$4,11 \times 10^{-7}$
<b>YFO50</b>	2,66	0,606	$3,53 \times 10^{-6}$
<b>YFO100</b>	2,50	0,337	$1,16 \times 10^{-6}$

**Figure 8.** The variation of band gap and activation energy with the amount of graphene oxide

### 3.3. Electrical characterization

The dielectric properties of the  $\text{YFeO}_3$  samples were investigated in the frequency range from 1 kHz to 5 MHz. The relative permittivity ( $\epsilon'$ ), dielectric loss ( $\epsilon''$ ) and alternating current conductivity ( $\sigma_{ac}$ ) were calculated using the following relations, respectively

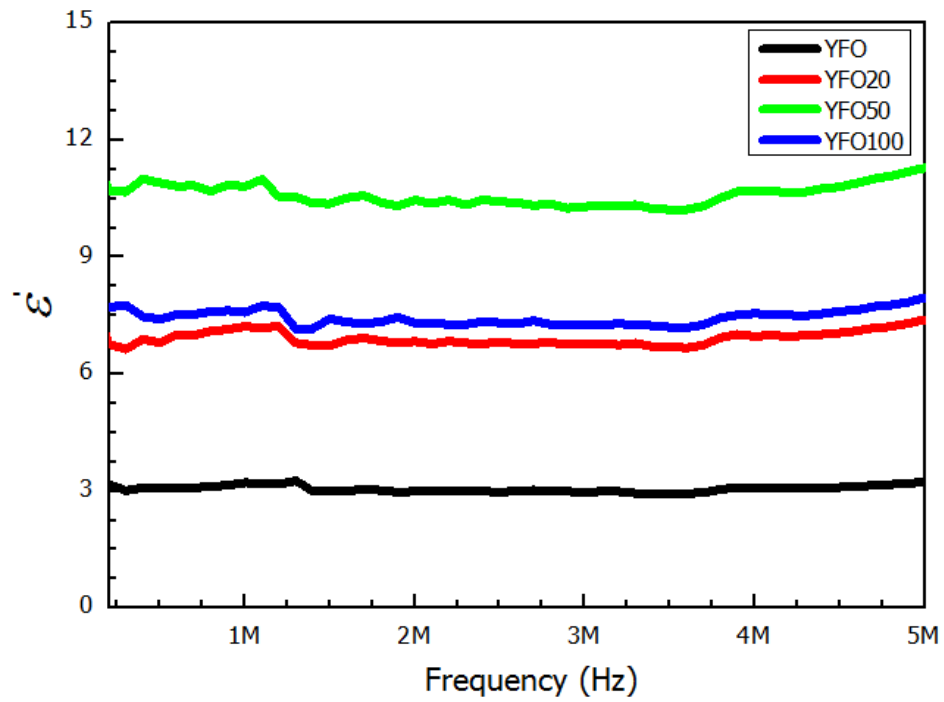
The relative permittivity ( $\epsilon'$ ), dielectric loss ( $\epsilon''$ ) and alternating current conductivity ( $\sigma_{ac}$ ) values of the samples were determined by the following relations [26, 27,36]:

$$\epsilon' = \frac{C_p \cdot t}{\epsilon_0 \cdot A} \quad (4)$$

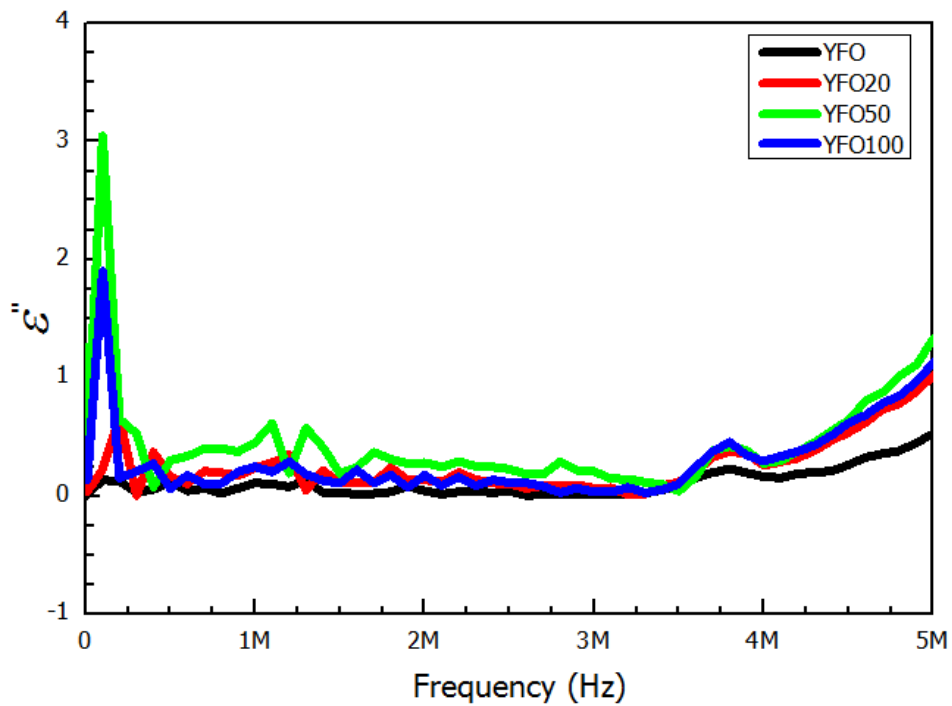
$$\epsilon'' = \tan \delta \cdot \epsilon' \quad (5)$$

$$\sigma_{ac} = \frac{l}{Z \cdot A} \quad (6)$$

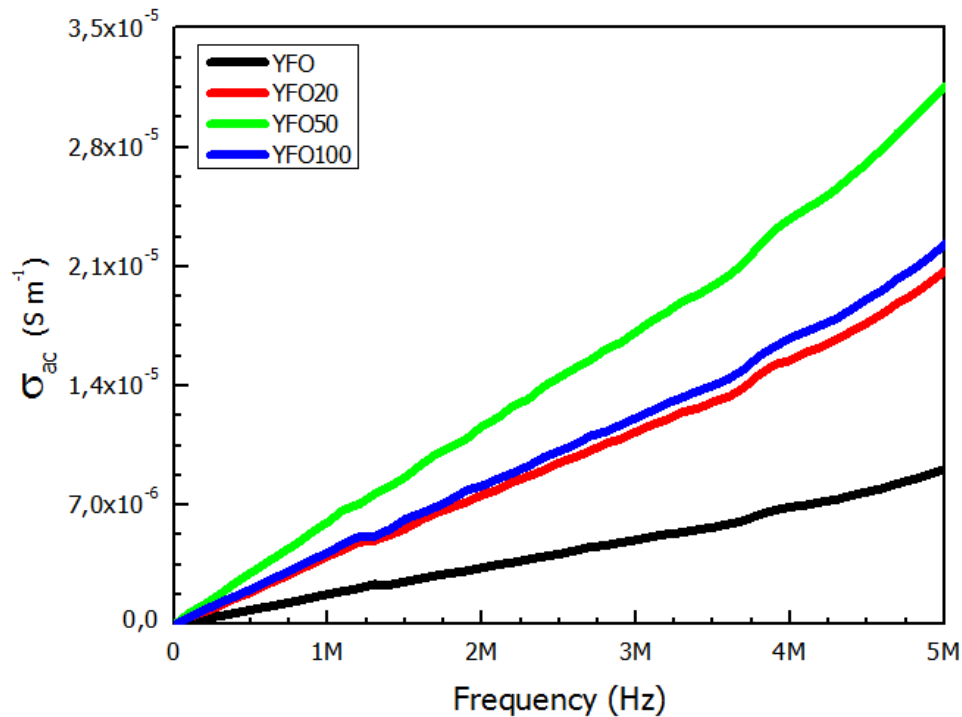
where  $\epsilon_0$  is the permittivity of free space,  $A$  is the area of the electrode,  $C$  is the capacitance,  $l$  is the thickness of the sample,  $\tan \delta$  is the loss tangent and  $Z$  is the impedance. Using the equations from 2 to 4, the relative permittivity ( $\epsilon'$ ), dielectric loss ( $\epsilon''$ ) and alternating current conductivity ( $\sigma_{ac}$ ) values were computed and their changes with increasing frequency were also plotted in Figs. 9, 10 and 11, respectively.



**Figure 9.** Relative permittivity as a function of frequency plots of the samples



**Figure 10.** Dielectric loss as a function of frequency plots of the samples



**Figure 11.** Alternating current conductivity vs. frequency plots of the samples

As can be seen from Figs. 9, 10 and 11, the relative permittivity ( $\epsilon'$ ), dielectric loss ( $\epsilon''$ ) and alternating current conductivity ( $\sigma_{ac}$ ) are significantly affected by graphene oxide content, respectively. While the values of  $\epsilon'$  and  $\epsilon''$  decrease with increasing of frequency, the  $\sigma_{ac}$  increases linearly, for each sample. In brief, graphene oxide-doping into  $\text{YFeO}_3$  affects its dielectric properties and alternating current conductivity values. The relative permittivity value for each sample decreases at the higher frequencies in comparison to the lower ones since the contribution of the dipolar polarization is small at the higher frequencies because of the inertia of the molecules and ions. At the same time, the orientation of dipoles cannot adapt to the higher applied alternating fields, and this leads to the dielectric loss. The increase in the dielectric permittivity with the different graphene oxide-amount can be a sign of a percolation transition from an insulator to conductor.

The relative permittivity values of the nano electroceramic samples at different frequencies were given in Table 3.

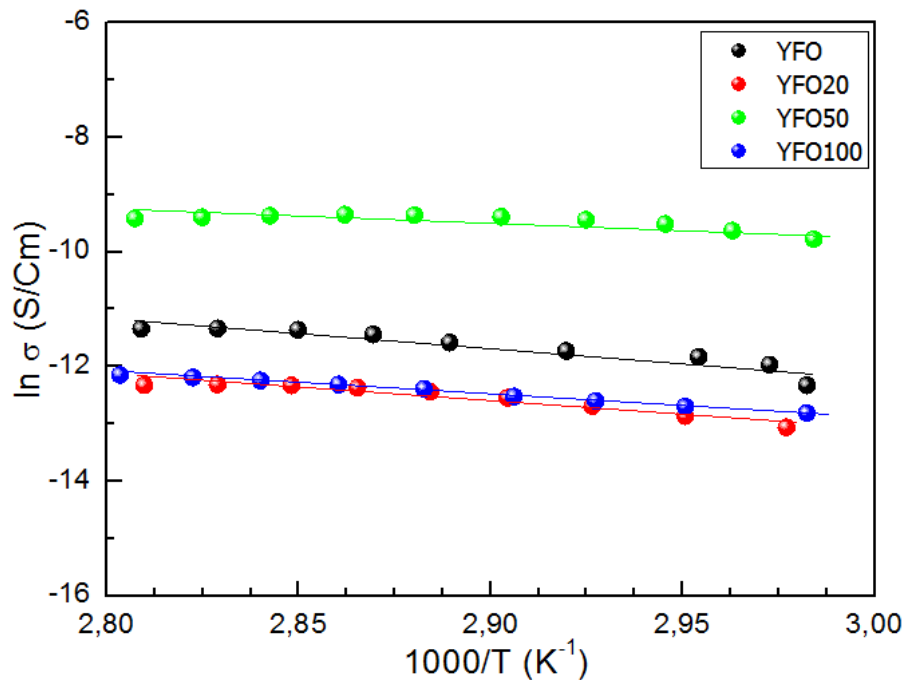
**Table 3.** The relative permittivity values of the samples at different frequencies

	$\epsilon'$								
	$\epsilon'_{1\text{kHz}}$	$\epsilon'_{10\text{kHz}}$	$\epsilon'_{100\text{kHz}}$	$\epsilon'_{500\text{kHz}}$	$\epsilon'_{1\text{MHz}}$	$\epsilon'_{2\text{MHz}}$	$\epsilon'_{3\text{MHz}}$	$\epsilon'_{4\text{MHz}}$	$\epsilon'_{5\text{MHz}}$
<b>YFO</b>	3,161	1,459	3,597	3,090	3,213	2,990	2,970	3,095	3,240
<b>YFO20</b>	7,330	3,349	8,037	6,797	7,219	6,833	6,776	6,966	7,391
<b>YFO50</b>	13,039	6,284	11,767	10,897	10,804	10,457	10,288	10,698	11,283
<b>YFO100</b>	8,575	3,349	7,400	7,406	7,580	7,307	7,263	7,551	7,957

In order to determine the temperature-dependent variation of the produced undoped and graphene oxide-doped nano electroceramics and their electrical transmission mechanism, the temperature-dependent conductivity curves were used. The conductivity curves were analyzed using the equation known as Arrhenius equation [19, 20, 24, 28, 29,37] :

$$\sigma_{dc} = \sigma_0 \cdot \exp\left(\frac{\Delta E_A}{kT}\right) \quad (7)$$

In Equation 7,  $\sigma_0$ ; is the exponential function,  $k$ ; Boltzmann constant and  $E_a$  represents the activation energy. In the light of the data,  $\ln\sigma-1000/T$  change graphs were drawn for all samples and activation energies corresponding to the donor level were calculated.



**Figure 12.** Temperature dependence of DC electrical conductivity of the yttrium ferrite samples

Fig.12. shows  $\ln\sigma-1000/T$  graph of the samples. As seen in Fig.12., electrical conductivity of the samples changed as the amount of graphene oxide dopant changed. The highest conductivity value was observed in YFO50 sample. It was also clearly seen that as the temperature increased, the conductivity increased. For all the samples, a single conductivity mechanism occurred in the measured temperature range. This conductivity mechanism formed by the thermal excitation of the carrier loads. As the temperature increased, the electrons were thermally stimulated to the conductivity band from the donor levels. With increasing temperature, more load carrier contributed to the electrical conductivity by exceeding the activation energy barrier. The activation energy of the samples was calculated from the slope of the lines in the graph and was given in Table 2. When the values in Table 2 were examined, activation energies decreased in parallel with the increasing graphene oxide ratio. YFO100 sample had the lowest activation energy value. In addition, the conductivity values of the samples at room temperature changed based on the graphene oxide dopant. Table 2 shows the conductivity values of the samples at room temperature.



#### 4. Discussion

Undoped and different mass ratio graphene oxide doped  $\text{YFeO}_3$  based nanocomposite samples were prepared via hydrothermal method. Structural, morphological, chemical, optical and electrical properties of the samples were investigated. All of the results obtained from characterization show a change with graphene oxide contribution. As a result, in this study is especially important in terms of, nanoceramic semiconducting materials, which have an important place in technological applications, can be produced especially at low cost and easily and also offering the possibility to dope of graphene oxide, which is characterized as the superior feature material of our time. In addition, this article is an important guide for future studies on researching of physical and chemical properties of  $\text{YFeO}_3$  based organic semiconductors with doping different elements and exploring the potential for use in advanced electronic applications.

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